Hello! My name is Drippy, and I’m here to tell you about mixtures. While this topic may seem simple, there are actually many different outcomes from mixing things together. Four, to be specific: solutions, suspensions, colloids, and emulsions. But don’t worry. With my help, identifying and understanding these four types will be easy as π!

**SOLUTION**
- Completely dissolved into the atomic level
- Can’t see the mixed particles with our eyes or a microscope.
- Can be filtered or separated by a physical change.

**Suspension**
- Mixture of solid particles dispersed in a liquid.
- Can be separated by filtering the mixture.

**Colloid**
- Mixture of solid particles dispersed in a liquid.
- May be transparent or cloudy.

**Emulsion**
- Type of colloid where dispersion is helped by an emulsifier.
- Forced mixtures.

**GOOD THINGS TO KNOW!**
- Solvent: the major part of the mixture that dissolves the solute.
- Solute: the smaller part of the mixture that gets dissolved.

Thank you to all the students, teachers and chemistry enthusiasts who entered the ACS ChemClubs and ChemMatters infographic contest! Check out the first of the four winners, as you learn more about the chemistry of mixtures. Congratulations to Aaron Herrera and Emerald Rawls, from Mapleton Expeditionary School of the Arts in Thornton, Colo. Watch for the other winners in the next three issues.

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